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SOLUTIONS MANUAL 3. Apply How many
atoms of hydrogen-1 are in a 1.007-g
sample? Recall that 1.007 amu is the
mass of one atom of hydrogen-1. Round
your answer to two significant digits.
1.007 g 1 H-1 atom $\underline{\quad}$ 1.672 10²⁴ g a.
6.023 10²³ H-1 atoms 4.

The MoleThe Mole

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Chapter 14 Solutions Manual CHAPTER
14 SOLUTIONS MANUAL 32. Explain the
similarities and differences between a 1
M solution of NaOH and a 1 m solution of
NaOH. Both solutions contain NaOH
(solute) dissolved in water (solvent). The
1 m solution contains 1 mole of NaOH
per kilogram of water; the 1M solution 3
_ 5 _ _ _ 1 _ _ _ _ _ (_

Mixtures and Solutions

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SOLUTIONS MANUAL 35. Aqueous
solutions of potassium iodide and silver
nitrate are mixed, forming the
precipitate silver iodide. chemical
equation: $KI(aq) + AgNO_3(aq) \rightarrow KNO_3(aq) + AgI(s)$
complete ionic equation: $K^+(aq) + I^-(aq) + Ag^+(aq) + NO_3^-(aq) \rightarrow K^+(aq) + NO_3^-(aq) + AgI(s)$

Chemical Reactions

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SOLUTIONS MANUAL 4. Determine why producing 1 million francium atoms per second is not enough to make measurements, such as density or boiling point. Even one million atoms collected together as a solid are microscopic. A grain of salt contains about 10^{15} sodium atoms.

The Periodic Table and Periodic Law

124 Chemistry: Matter and Change • Chapter 8 Solutions Manual 45. O_3 O_2 O O O O 46. Challenge Draw the Lewis resonance structure for the ion SO_3^{2-} . SO_2 SO_3^{2-} SO_2 SO_3^{2-} Draw the expanded octet Lewis structure of each molecule. 47. ClF_3 ClF_4 ClF_5 48. PCl_5 PCl_6^- 49. Challenge Draw the Lewis structure for the molecule formed when six fluorine atoms and

Covalent Bonding

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